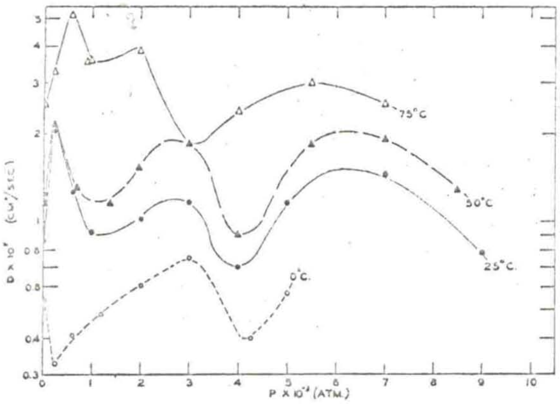

FIG. 1. Diffusion coefficients *versus* pressure, 0.1 N $Hg(NO_3)_2$.

minimum shifts toward lower pressure as the temperature increases. The curves also show a maximum in the pressure range 250–600 atmospheres which moves to

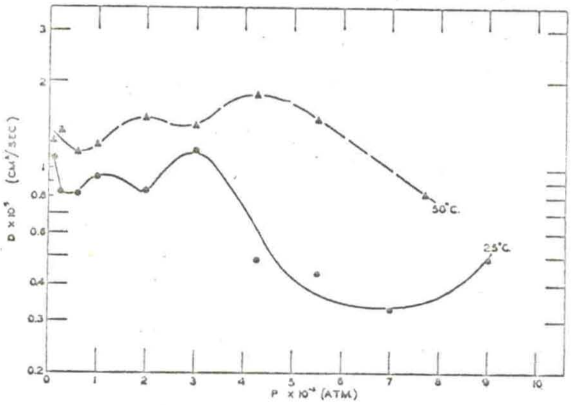

FIG. 2. Diffusion coefficients *versus* pressure, 0.1 N $CaCl_2$.

lower pressure as temperature decreases and disappears at 0°C.

Curves of activation enthalpy and the increase in

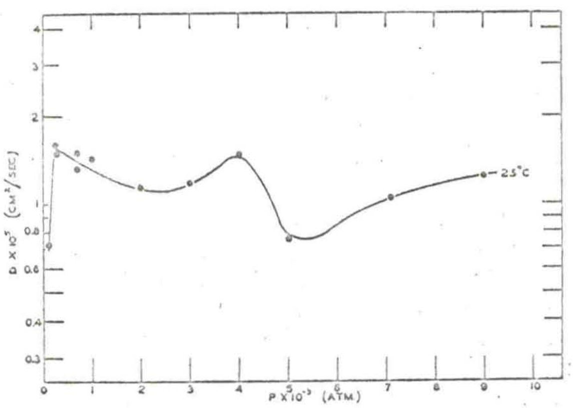

FIG. 3. Diffusion coefficients *versus* pressure, 0.1 N $Ca(NO_3)_2$.

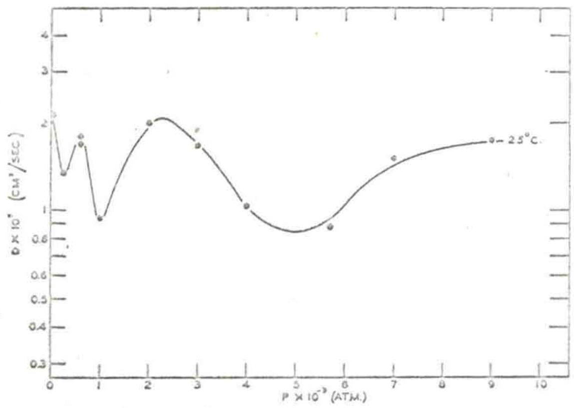

FIG. 4. Diffusion coefficients *versus* pressure, 0.1 N $TlNO_3$.

activation entropy from one atmosphere to P are shown in Figs. 7 and 8. Any interpretation of these curves

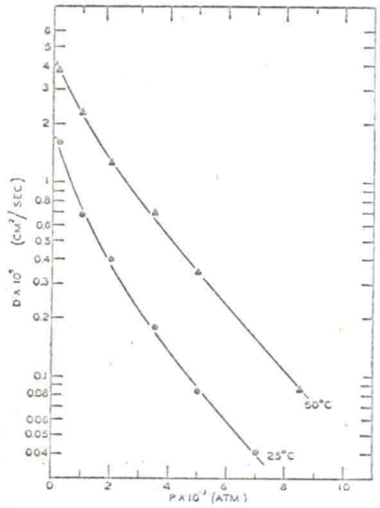

FIG. 5. Diffusion coefficients *versus* pressure, 0.01 M $HgCl_2$ in n-butanol.

would be dubious because the activation enthalpy and entropy depend on the displacement of two or more

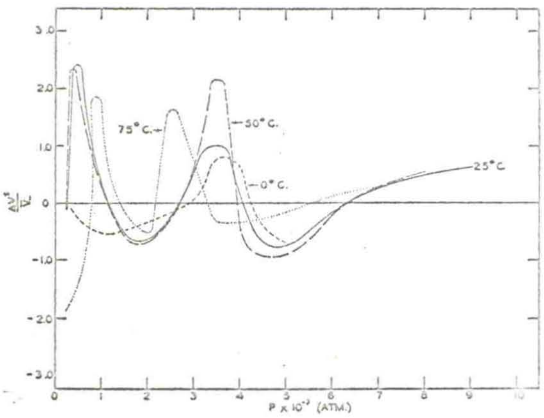

FIG. 6. Activation volume ratios *versus* pressure. 0.1 N $Hg(NO_3)_2$.